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## PROBING BRØNSTED ACIDITY OF PROTONIC ZEOLITES WITH VARIABLE-TEMPERATURE INFRARED SPECTROSCOPY

UDC 539

*Most industrial applications of zeolites as solid-acid catalysts rely on their high Brønsted acidity, which affects both catalytic activity and selectivity, and hence the convenience to find an accurate experimental technique for measuring the acid strength. The enthalpy change,  $\Delta H^0$ , involved in the hydrogen bonding interaction between a weak base (such as carbon monoxide) and the Brønsted acid [Si(OH)Al] hydroxyl groups should correlate directly with the zeolite acid strength. However, on account of simplicity, the bathochromic shift of the O–H stretching frequency,  $\Delta\nu(\text{OH})$ , is usually measured by IR spectroscopy at a (fixed) low temperature instead of  $\Delta H^0$  and correlated with the acid strength for ranking the zeolite acidity. Herein, the use of variable-temperature IR spectroscopy to determine simultaneously  $\Delta H^0$  and  $\Delta\nu(\text{OH})$  is demonstrated, followed by a review of recent experimental results showing that the practice of ranking the acid strength by the corresponding O–H frequency shift probed by a weak base could be misleading; and that can be so much the case of zeolites showing a wide range of structure types.*

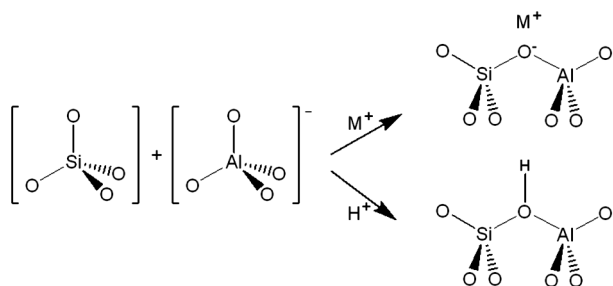
*Keywords:* Brønsted acidity, infrared spectroscopy, zeolites.

### 1. Introduction

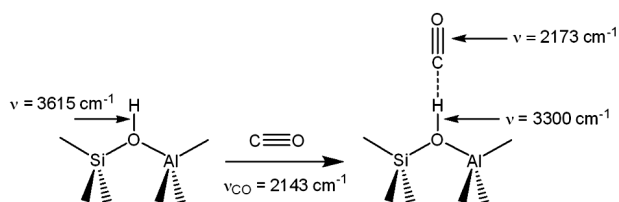
Zeolites are three-dimensional aluminosilicates composed of corner sharing tetrahedral  $\text{SiO}_4$  and  $\text{AlO}_4$  units that form a crystalline framework characteristic of each zeolite structure type, which encloses void channels and cavities having entrances of about 3 to 8 Å in diameter [1]. The replacement of Al for Si in the tetrahedral units generates a net negative charge that has to be balanced by extraframework (exchangeable) cations or protons, as shown in Scheme 1; in the latter case, protonic zeolites, which show a distinctive Brønsted acidity, are generated. This acidity endows protonic zeolites with widespread applications (as solid acid catalysts) in a large range of chemical processes, which span the petrochemical industry, oil refining, methanol to olefin (MTO) conversion, and the production of fine chemicals, to name only a few examples [2–8]. The strength of their catalytic [Si(OH)Al] hydroxyl groups constitutes a main factor determining the catalytic activity and the se-

lectivity of protonic zeolites; and, hence, the importance of having a reliable method to quantify relative acidity. For that purpose, several instrumental techniques can (in principle) be used; among them, solid-state NMR spectroscopy, calorimetry of an adsorbed base (such as  $\text{NH}_3$  or pyridine), and temperature programmed desorption [9–16].

The aforementioned techniques aside, infrared spectroscopy (at liquid nitrogen temperature) of an adsorbed probe molecule (a weak base) is currently the method most commonly used to rank the zeolite acidity, carbon monoxide often being the probe molecule of choice [17–21], although dinitrogen or carbon dioxide could also be used [22–24]. As shown in Scheme 2, CO interacts with the Brønsted-acid hydroxyl group of the zeolite forming a hydrogen-bonded  $\text{OH} \cdots \text{CO}$  complex, which gives rise to a large low-frequency shift,  $\Delta\nu(\text{OH})$ , of the O–H stretching mode [25, 26] that can easily be monitored by IR spectroscopy, and the corresponding magnitude of  $\Delta\nu(\text{OH})$  is often used to rank the Brønsted acidity of zeolites, assuming that such a magnitude is proportional to the corresponding hydrogen-



Scheme 1

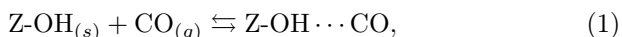


Scheme 2. Wavenumbers are orientative. Actual values depend on the specific zeolite being considered

bond strength. Nevertheless, recent experimental results obtained by means of variable-temperature IR (VTIR) spectroscopy have shown that such an assumption can be misleading; particularly when dealing with zeolites having different structure types [27, 28]. The current state of knowledge in this field is reviewed below, after a brief introduction to the physical tenets and practical implementation of the VTIR method.

## 2. Outline of the Variable-Temperature IR (VTIR) Spectroscopic Method

The VTIR method, recently developed, is an instrumental technique particularly well suited to gain access to gas-solid physisorption thermodynamics, while simultaneously obtaining the spectroscopic signature of the adsorption complex, provided that either the solid adsorbent or the molecule being adsorbed from the gas phase has an IR-active mode that undergoes a change brought about by the adsorption process. Referring specifically to the adsorption of CO on protonic zeolites (Sch. 2), let Eq. (1) below represent the adsorption equilibrium:



where Z-OH stands for the Brønsted-acid hydroxyl group of the zeolite, and Z-OH  $\cdots$  CO is the hydrogen-bonded adsorption complex. In this case, the forma-

tion of the adsorption complex involves a characteristic change in the stretching frequency of both, the zeolite hydroxyl group and the (adsorbed) CO molecule; and, hence, any of the corresponding IR absorption bands could (in principle) be used to monitor the adsorption process; however, most frequently the band coming from the O-H stretching mode of the hydroxyl group is the one chosen for the experimental measurements. The integrated intensity of that IR absorption band should be proportional to the surface coverage,  $\vartheta$ , according to the Lambert–Beer law. Therefore, it gives information on the activity (in the thermodynamic sense) of both, the adsorbed species and the empty adsorption sites. Simultaneously, the equilibrium pressure,  $p$ , gives the activity of the gas phase. Hence, by measuring the IR absorbance and equilibrium pressure at any given temperature, the equilibrium constant,  $K$ , of the adsorption process (at that temperature) can be determined.

Assuming that changes in specific heat are negligible, the variation of  $K$  with temperature,  $T$ , is related to the standard adsorption enthalpy,  $\Delta H^0$ , and entropy,  $\Delta S^0$ , through the well-known van't Hoff equation:

$$K(T) = \exp(-\Delta H^0/RT) \exp(\Delta S^0/R). \quad (2)$$

For a Langmuir-type adsorption process, combination of Eq. (2) with the Langmuir equation (3) leads to Eq. (4) below:

$$\vartheta = K(T)p/[1 + K(T)p], \quad (3)$$

$$\ln\{\vartheta/[(1 - \vartheta)p]\} = (-\Delta H^0/RT) + (\Delta S^0/R). \quad (4)$$

Alternatively, Eq. (4) can be written as

$$\ln\{A/[(A_M - A)p]\} = (-\Delta H^0/RT) + (\Delta S^0/R), \quad (5)$$

where  $A$  is the actual IR absorbance being measured, and  $A_M$  stands for the maximum absorbance, which corresponds to  $\vartheta = 1$ . It is clear that, after determining  $\vartheta$  (or relative IR absorbance) as a function of  $T$  and  $p$  over a temperature range, Eq. (4) or (5) provides the direct access to the corresponding values of  $\Delta H^0$  and  $\Delta S^0$  which characterize the adsorption equilibrium process. It is worth of note, however, that the assumption made that  $\Delta H^0$  and  $\Delta S^0$  are both temperature-independent implies that  $\Delta C_p^0$  (the difference of specific heats at a constant pressure between the gas phase and the adsorbed phase)

is nil. Strictly speaking, this condition is never fulfilled, because the degrees of freedom of the adsorbate molecule are not the same in both states. However, translational degrees of freedom in the gas phase are replaced by low-lying vibrational modes of the adsorbed molecule [29], which also retains some rotational freedom. Hence,  $\Delta C_p^0$  is expected to be much smaller than  $3R$ , and the inclusion of the corresponding correction would not affect significantly the final results. Nevertheless, it should be kept in mind that both  $\Delta H^0$  and  $\Delta S^0$  are the values of the thermodynamic quantities actually temperature-dependent and averaged over the temperature range spanned by the VTIR spectroscopic measurement. It is therefore appropriate to refer both of them to the average of that temperature. The further discussion on the potential use and limitations of the VTIR method, including reasons for adopting the Langmuir-type adsorption model, can be found elsewhere [30].

On the experimental side, a properly designed IR cell is needed for recording VTIR spectra. Some commercial cells are adaptable for such a purpose, but most of the experimental results discussed below were obtained by using a homemade IR cell described in detail elsewhere [31]. Measurements are run by dosing a fixed amount of the adsorbate (CO) into the liquid nitrogen cooled cell, which contains a previously activated (outgassed) solid wafer. After that, the cell is closed, and a series of IR spectra is obtained at an increasing temperature, while simultaneously recording the temperature and equilibrium pressure. The cell is thus operated as a closed system in the thermodynamic sense; which is at variance with calorimetric or volumetric gas adsorption measurements usually performed in open systems. Further details on the cell operation and the sample preparation are given in Section 3.1.

### 3. Selected Case Studies

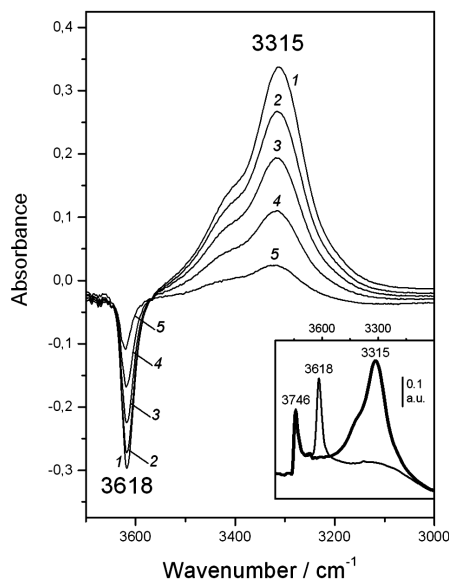
#### 3.1. Zeolite H-ZSM-5: Structure-type MFI

A sample of H-ZSM-5, having a nominal Si:Al ratio of 30:1, was obtained from a commercial firm and checked by powder X-ray diffraction, which showed a good crystallinity and the absence of any diffraction line not assignable to the corresponding MFI structure type. For VTIR spectroscopy, a thin self-supported wafer of zeolite was prepared and outgassed by the thermal treatment in a dynamic vacuum (residual pressure smaller than  $10^{-4}$  mbar) for

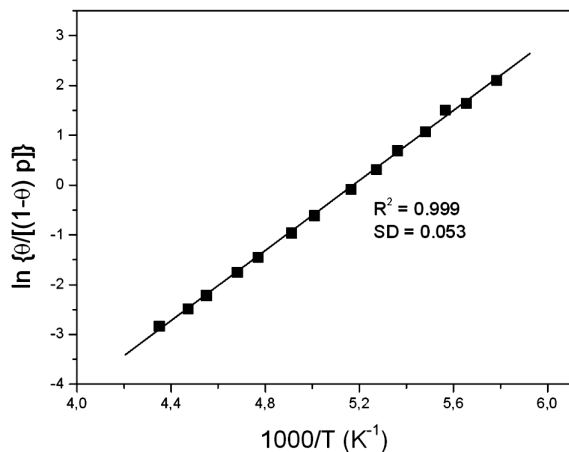
5 h at 670 K inside an IR cell [31] that allowed the *in situ* sample treatment, gas dosage, and variable-temperature IR spectroscopy to be carried out. After the thermal treatment, the cell was cooled with liquid nitrogen, and 0.2 mbar of helium was admitted into the sample compartment in order to improve the thermal contact between the cell body and the sample wafer. After recording the zeolite blank spectrum, the IR cell was dosed with a fixed amount of CO (just enough to form a 1:1 adsorption complex on every Brønsted-acid hydroxyl group of zeolite) and closed, and a series of IR spectra was recorded upon the gradual warming up of the cell, while simultaneously registering the temperature and gas equilibrium pressure. For that purpose, the cell was equipped with a platinum resistance thermometer (Tinsley) and a capacitance pressure gauge (MKS, Baratron). The precision of these measurements was better than  $\pm 10^{-2}$  mbar and  $\pm 2$  K for the pressure and temperature, respectively. The pressure correction (for helium) was determined from a previous calibration plot. Transmission FT-IR spectra were recorded at  $2\text{ cm}^{-1}$  resolution on a Bruker IFS66 instrument, accumulating 64 scans for each spectrum. To avoid the repetition, please notice that the same procedure was used for recording the VTIR spectra of the other zeolite samples described below.

As shown in the inset to Fig. 1, the blank spectrum of zeolite displays (in the O–H stretching region) the characteristic IR absorption bands at 3746 and  $3618\text{ cm}^{-1}$ , which correspond, respectively, to silanols and to the bridged [Si(OH)Al] groups that constitute the zeolite Brønsted-acid sites [17, 19]. Upon the interaction with CO (in the low pressure range studied herein), the silanol band was not altered, but that at  $3618\text{ cm}^{-1}$  was eroded, and a new (and much broader) band appeared simultaneously at  $3315\text{ cm}^{-1}$ , thus testifying to the formation of the corresponding OH $\cdots$ CO hydrogen-bonded complex (being  $\Delta\nu(\text{OH}) = 303\text{ cm}^{-1}$ ).

Figure 1 (main body) displays representative VTIR spectra, which are shown in the difference mode (*i.e.*, after subtracting the blank spectrum of zeolite) in order to avoid any spurious contribution from the background. They show the expected intensity variation of the band at  $3618\text{ cm}^{-1}$  (bridging [Si(OH)Al] hydroxyl groups), as the temperature is raised, which parallels that shown by the band at  $3315\text{ cm}^{-1}$  (O–H stretching mode of the hydrogen-bonded OH $\cdots$ CO

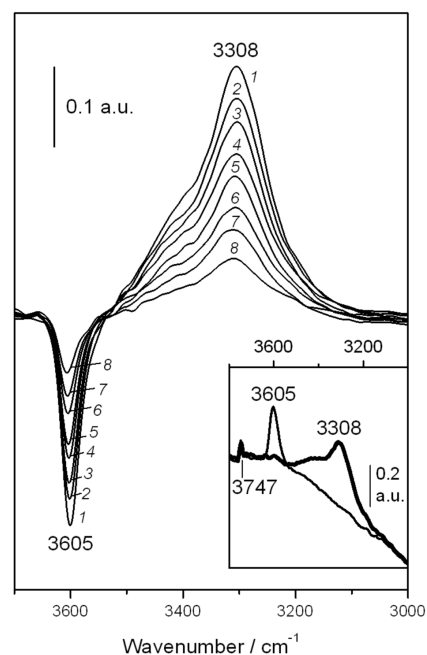


**Fig. 1.** Representative variable-temperature IR spectra (O–H stretching region) of CO adsorbed on H-ZSM-5. The spectra are shown in the difference mode (zeolite blank subtracted). Temperature (in K) and equilibrium pressure (mbar, in brackets) are as follows: 1 – 173 (4.82); 2 – 190 (6.23); 3 – 202 (6.98); 4 – 214 (7.64); 5 – 230 (8.39). Inset shows the blank zeolite spectrum and the effect of dosing with CO at 77 K (bold line)



**Fig. 2.** Plot of the left-hand side of Eq. (4) against the reciprocal temperature for CO adsorbed on H-ZSM-5; data obtained from the O–H stretching band at 3618  $\text{cm}^{-1}$ . R, linear regression coefficient; SD, standard deviation

complex). From the integrated intensity of the O–H stretching band at 3618  $\text{cm}^{-1}$  measured for the whole set of VTIR spectra recorded, the corresponding values for the coverage of adsorption sites,  $\vartheta$ , were de-



**Fig. 3.** As in Fig. 1; CO adsorbed on H-FER. Temperature (in K) and equilibrium pressure (mbar, in brackets) are as follows: 1 – 167 (0.76); 2 – 178 (1.07); 3 – 133 (1.25); 4 – 191 (1.48); 5 – 195 (1.59); 6 – 202 (1.80); 7 – 206 (1.92); 8 – 215 (2.15)

duced. Note that the intensity of the O–H stretching band divided by its maximum value (*i.e.*, that corresponding to the blank zeolite spectrum) gives directly the fraction,  $1 - \vartheta$ , of free OH sites. From the whole set of  $\vartheta$ ,  $T$  and  $p$  values, the plot of Eq. (4) shown in Fig. 2 was obtained. An excellent linear fit can be seen, which led to the values of  $\Delta H^0 = -29 \text{ kJ mol}^{-1}$  (and  $\Delta S^0 = -150 \text{ J mol}^{-1}\text{K}^{-1}$ ) for the standard enthalpy (and entropy) of the formation of the  $\text{OH}\cdots\text{CO}$  hydrogen-bonded complex. The estimated error limits are  $\pm 2 \text{ kJ mol}^{-1}$  for the enthalpy and  $\pm 10 \text{ J mol}^{-1}\text{K}^{-1}$  for the entropy.

### 3.2. Protonic ferrierite (H-FER): Structure-type FER

The ferrierite sample was obtained from Zeolyst International, it was in the ammonium form ( $\text{NH}_4\text{-FER}$ ) and had a Si:Al ratio of 27.5:1; the conversion into the protonic form (H-FER) was achieved and checked, as described elsewhere [32]. The IR spectrum of a blank zeolite wafer (depicted in the inset of Fig. 3) showed the characteristic O–H stretching band of the Brønsted-acid hydroxyl groups at 3605  $\text{cm}^{-1}$  together

with a weaker silanol band at  $3747\text{ cm}^{-1}$  (which is of no concern herein).

Representative IR spectra of carbon monoxide adsorbed at a variable temperature on a thin self-supported zeolite wafer are shown in Fig. 3 (main body). As expected, the interaction of zeolite with adsorbed CO left the silanol band unaltered, while the intensity of the band at  $3605\text{ cm}^{-1}$  decreased to an extent depending on the temperature. Simultaneously, a new (broader) IR absorption band due to  $\text{OH}\cdots\text{CO}$  species developed at  $3308\text{ cm}^{-1}$ . After determining the integrated intensity of the whole set of bands at  $3605\text{ cm}^{-1}$ , the van't Hoff plot depicted in Fig. 4 was obtained. From that linear plot, the value of  $\Delta H^0 = -28.6\text{ kJ mol}^{-1}$  was derived for the standard enthalpy of the H-bond formation between adsorbed CO and the H-FER Brønsted acid sites. The corresponding standard entropy change resulted to be  $\Delta S^0 = -147\text{ J mol}^{-1}\text{ K}^{-1}$ .

### 3.3. H-MCM-22: Structure-type MWW

From the parent material, MCM-29P [33, 34], the protonic zeolite H-MCM-22 (Si:Al ratio 24.5:1) was prepared as described elsewhere [28] and checked by powder X-ray diffraction, which confirmed a good crystallinity and the absence of any diffraction line not assignable to the corresponding MWW structure type. The Si:Al ratio was determined by X-ray fluorescence analysis.

The FT-IR spectrum of zeolite in the O-H stretching region is displayed in Fig. 5 (inset). It shows the characteristic IR absorption band of Brønsted-acid hydroxyl groups at  $3625\text{ cm}^{-1}$  (and the silanol band at  $3750\text{ cm}^{-1}$ ). Dosing with CO (at 77 K) was found to erode the  $3625\text{ cm}^{-1}$  band and to bring about the corresponding IR absorption band of the hydrogen-bonded  $\text{OH}\cdots\text{CO}$  complex, which peaks at  $3305\text{ cm}^{-1}$ , thus showing a  $\Delta\nu(\text{OH})$  shift of  $-320\text{ cm}^{-1}$ . Representative variable-temperature IR spectra are shown in the main body of Fig. 5. After measuring the integrated intensity of the bands at  $3625\text{ cm}^{-1}$  for the whole set of VTIR spectra recorded, the van't Hoff plot shown in Fig. 6 was obtained, from which the standard enthalpy of the formation of hydrogen-bonded complexes resulted to be  $\Delta H^0 = -23.5\text{ kJ mol}^{-1}$ . The corresponding value of  $\Delta S^0$  was  $-130\text{ J mol}^{-1}\text{ K}^{-1}$ . As a further check, the interaction enthalpy of CO with the same H-MCM-22 sample was also determined by adsorption

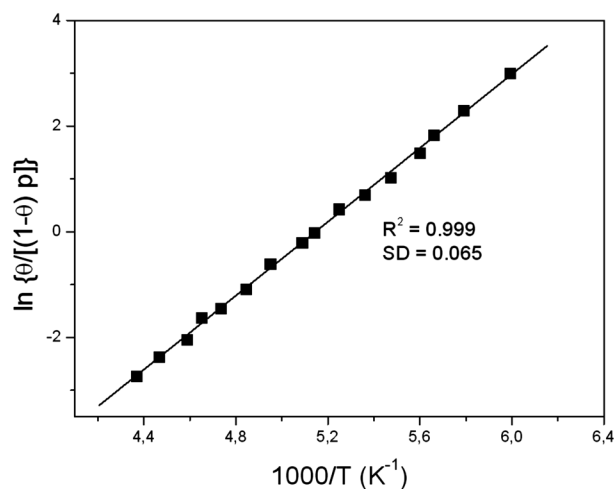


Fig. 4. Plot of the left-hand side of Eq. (4) against the reciprocal temperature for CO adsorbed on H-FER; data obtained from the O-H stretching band at  $3605\text{ cm}^{-1}$

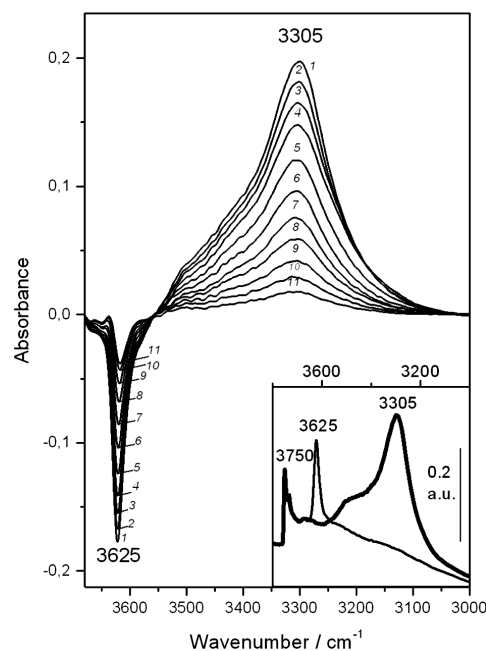
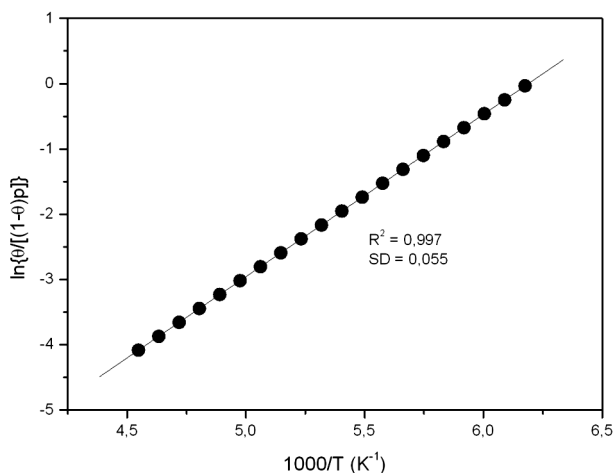
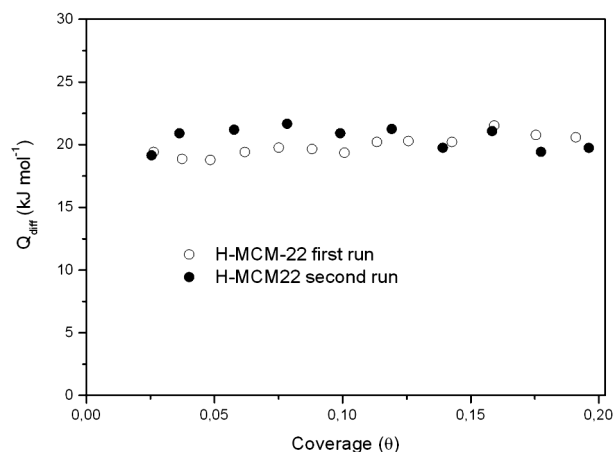


Fig. 5. As in Fig. 1; CO adsorbed on H-MCM-22. Temperature (in K) and equilibrium pressure (mbar, in brackets) are as follows: 1 – 154 (6.52); 2 – 160 (6.99); 3 – 166 (7.37); 4 – 172 (7.72); 5 – 178 (8.01); 6 – 184 (8.29); 7 – 190 (8.55); 8 – 196 (8.75); 9 – 202 (8.93); 10 – 208 (9.11); 11 – 214 (9.24)

calorimetry at 303 K (see Ref. [35] for details). Figure 7 shows the obtained plots of the differential heat of adsorption as a function of the coverage. Note that, in order to check the reproducibility, the sec-



**Fig. 6.** Plot of the left-hand side of Eq. (4) against the reciprocal temperature for CO adsorbed on H-MCM-22; data obtained from the O–H stretching band at 3625  $\text{cm}^{-1}$

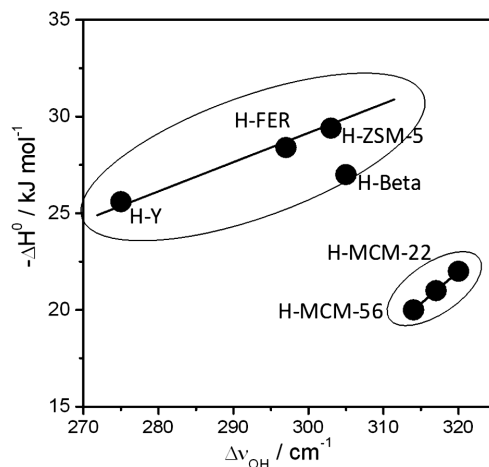


**Fig. 7.** Adsorption heat (as a function of the coverage) of CO on H-MCM-22, measured by calorimetry at 303 K. Two independent series of measurements are shown

ond measurement run was performed after outgassing CO adsorbed in the first run. The average of both sets of measurements gave an adsorption heat of  $21(\pm 2)$   $\text{kJ mol}^{-1}$ , to be compared with the value of  $\Delta H^0 = -23.5(\pm 2)$   $\text{kJ mol}^{-1}$  obtained by VTIR spectroscopy; both values are approximately equal within the limits of experimental errors.

#### 4. Discussion and Conclusions

The foregoing case studies demonstrate how VTIR spectroscopy can be used to quantify the Brønsted acid strength of protonic zeolites. Moreover, the case of CO adsorption on H-MCM-22 (Section 3.3) shows



**Fig. 8.** Correlation between  $\Delta H^0$  and  $\Delta\nu(\text{OH})$  for the CO hydrogen bonding in protonic zeolites

#### Relevant experimental data for the CO hydrogen bonding in protonic zeolites

Zeolite	Structure type	Si/Al ratio	$\Delta\nu_{\text{OH}}^a$ , $\text{cm}^{-1}$	$\Delta H^b$ , $\text{kJ mol}^{-1}$
H-Beta	BEA	20	305	27
H-ZSM-5	MFI	30	303	29
H-FER	FER	27.5	297	28.6
H-Y	FAU	5.6	275	25.6
H-MCM-22	MWW	24.5	320	23.5
H-MCM-22	MWW	16.4	317	21
H-MCM-56	MWW	16	316	20

<sup>a</sup> Red-shift of the Brønsted-acid OH group upon hydrogen bonding with CO.

<sup>b</sup> Standard enthalpy change in the formation of the  $\text{OH} \cdots \text{CO}$  complex.

that the  $\Delta H^0$  value obtained by VTIR spectroscopy agrees (within experimental error) with the corresponding adsorption heat determined by calorimetry. However, it should be remarked that, besides probing the interaction energy, VTIR spectroscopy shows simultaneously the nature of the gas-solid adsorption complex, which is characterized by the corresponding IR spectrum, thus providing the site-specific information that is not made available by calorimetry; unless the latter technique is coupled with IR spectroscopic measurements. The fact of being site-specific confers a clear advantage to VTIR spectroscopy over adsorption calorimetry when dealing with zeolites containing point defects (such as, for instance, extra-framework aluminium species) that can adsorb CO [36, 37]. In general terms, CO adsorption

complexes formed on defective surface sites would be clearly discriminated (on account of their different spectroscopic signature) from H-bonded adsorption complexes when probed by VTIR spectroscopy; but that could not be the case when using adsorption calorimetry instead.

In addition to the foregoing case studies, VTIR spectroscopy has also been used to probe the interaction of CO with several other protonic zeolites, namely, H-Beta [27], H-Y [38], H-MCM-22 (Si/Al = 16) and H-MCM-56 [35]. The whole set of results obtained is summarized in the table, while the derived correlation between  $\Delta H^0$  and  $\Delta\nu(\text{OH})$  is depicted in Fig. 8. Clearly, zeolites H-Y, H-FER and H-ZSM-5 do show the expected (linear) correlation between the corresponding (absolute) values of  $\Delta\nu(\text{OH})$  and  $\Delta H^0$ , but H-Beta displays a significant deviation; and the three MWW structure-type zeolites come wide apart from the others. These findings strongly suggest that the rather common practice of ranking the Brønsted acid strength of protonic zeolites by the corresponding value of  $\Delta\nu(\text{OH})$ , probed by IR spectroscopy (at a fixed temperature) of adsorbed carbon monoxide, should be viewed with some caution, especially when dealing with widely different structure types. To conclude, it seems pertinent to remark that, some time ago, Tsai *et al.* [39] and Zhou *et al.* [40] have found that H-MCM-22 showed a weaker Brønsted acidity than H-ZSM-5, when both zeolites were probed by the temperature programmed desorption of ammonia; and yet  $\Delta\nu(\text{OH})$  is significantly larger for H-MCM-22 (Table). These facts give a further support to the foregoing arguments against the indiscriminated use of  $\Delta\nu(\text{OH})$  to rank the Brønsted-acid strength of protonic zeolites.

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ВИЗНАЧЕННЯ КИСЛОТНОСТІ  
БРОНСТЕДА ДЛЯ ПРОТОНОВАНИХ ЦЕОЛІТІВ  
ІЧ-СПЕКТРОСКОПІЄЮ ЗА ЗМІННОЇ ТЕМПЕРАТУРИ  
Р е з ю м е

Багато застосувань цеолітів у промисловості як кислотних каталізаторів у твердій фазі ґрунтуються на їх високій кислотності Бронстеда, що визначає їх каталітичні властивості і селективність і на наявності точної апаратури для вимірювання кислотності. Зміна ентальпії  $\Delta H^0$ , що супутня взаємодії за допомогою водневого зв'язку між слабкою основою (такою як окис вуглецю) і гідроксильною групою в [Si(OH)Al] з кислотністю Бронстеда, має безпосередньо корелювати з кислотністю цеолітів. Для простоти замість  $\Delta H^0$  зазвичай вимірюється при низькій (фіксованій) температурі з застосуванням ІЧ-спектроскопії багочромний зсув частоти коливань OH,  $\Delta\nu(\text{OH})$ , що корелює з силою кислоти для різних кислотностей цеолітів. У даній роботі ІЧ-спектроскопію за змінної температури застосовано для одночасного визначення  $\Delta H^0$  і  $\Delta\nu(\text{OH})$  та дано огляд результатів недавніх експериментів. Показано, що практика визначення сили кислоти по зрушенню частоти O–H коливань при взаємодії зі слабкою основою може бути помилковою, особливо при порівнянні цеолітів з широким діапазоном структур.